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Handbook Of Chemical Engineering Calculations

Hicks is the author of more than 20 books in engineering and related fields, including Handbook of Energy Engineering Calculations and Civil Engineering Formulas, Second Edition. Nicholas Chohey (deceased) was Editor-in-Chief at Chemical Engineering magazine, with more than 45 years' experience in engineering and publishing.

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Hicks holds a bachelor's degree in Mechanical Engineering from Cooper Union School of Engineering in New York. He is the author or coauthor of more than 20 books in engineering and related fields, including Civil Engineering Formulas, Handbook of

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(Perry and Chilton—Chemical Engineers' Handbook, McGraw-Hill, 1973.)
1.9 HEAT CAPACITY OF REAL GASES
Calculate the heat capacity C_p of ethane vapor at 400 K (260 F) and 50 atm (5065 kPa). Also estimate the heat-capacity ratio C_p/C_v at these conditions.

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1.1 MOLAR GAS CONSTANT
1.2 ESTIMATION OF CRITICAL TEMPERATURE FROM EMPIRICAL CORRELATION
1.2
1.3 CRITICAL PROPERTIES FROM GROUP-CONTRIBUTION METHOD
1.3
1.4 REDLICH-KWONG EQUATION OF STATE
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1.8

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The following equations can be used to predict heat-transfer coefficients for condensation on banks of horizontal tubes: For laminar-film condensation, $\rho L^2 g n L^{1/3} \frac{1}{1/6} \mu W N r h c = a k$ where $a = 0.951$ for triangular tube patterns, 0.904 for rotated square tube patterns, or 0.856 for square tube patterns.

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Calculation Procedure 1. Assume a basis. Assume gas is at standard conditions, that is, 1 g·mol gas at 1 atm (101.3 kPa) pressure and 0 C (273 K, or 492 R), occupying a volume of 22.4 L. 2. Compute the gas constant. Apply suitable conversion factors and obtain the gas constant in various units. Use $PV = RT$; that is, $R = PV/T$. Thus, a.

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